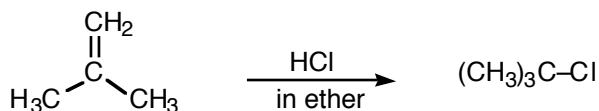
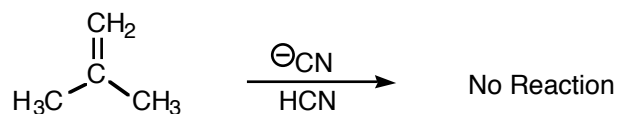
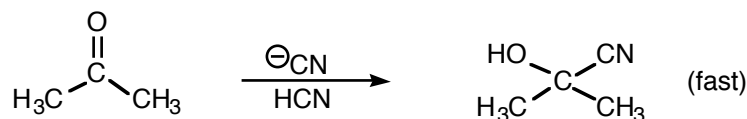


WORKSHOP, Chapter 17

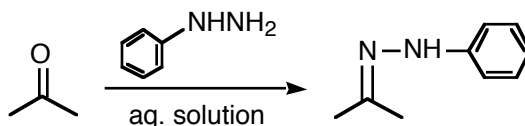
Carbonyl Compounds: Nucleophilic Addition

1. a. Discuss the similarities and differences in the structures and reactivities of the double bonds of 2-methylpropene and 2-propanone.
- b. With part a in mind, account for the following observations. Identify the nucleophile and the electrophile for the first step of each reaction.

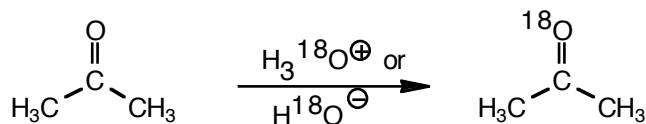


2. Write detailed electron-pushing mechanisms that explain the stated observations.

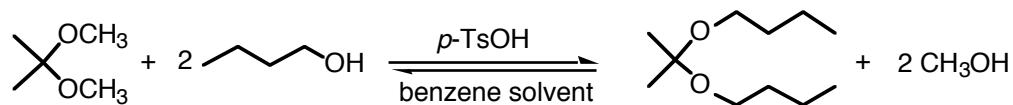
- a. The following reaction occurs rapidly at pH = 5 but fails at pH = 1.



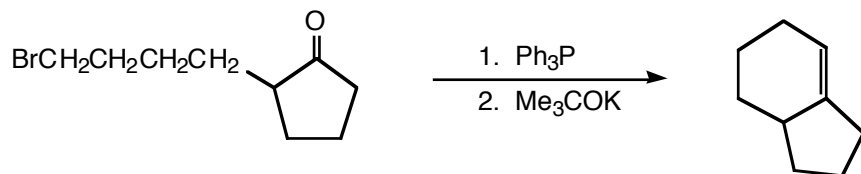
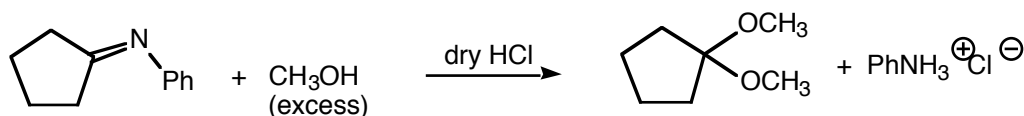
- b. Acetone can be labeled with ^{18}O either acidic or basic solution. The labeling of acetone with ^{18}O is catalyzed by both acid and base.



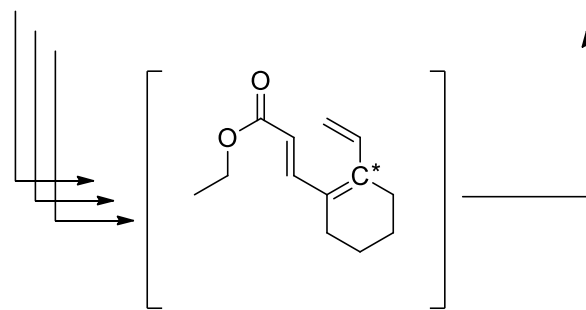
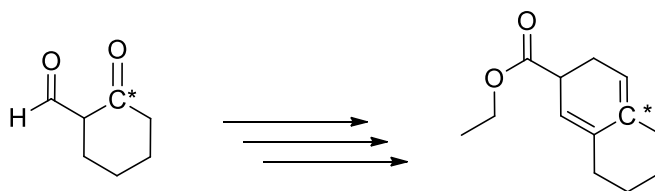
3. When 2,2-dimethoxypropane is mixed with *n*-butanol and some *p*-toluenesulfonic acid in benzene, the equilibrium shown below is established. The lowest boiling component of this mixture is methanol. Provide a reasonable mechanism for the reaction and suggest an experimental technique that could be used to shift the equilibrium to the right and thus make the reaction useful for synthesizing the dibutylketal. (Note: There is no water present. Therefore, the reaction does **NOT** proceed by hydrolysis of the ketal to acetone.)



4. Give a reasonable mechanism for each of the following reactions, clearly showing all important intermediates. Show movement of electron pairs with curved arrows.



5. Try this and make this ester from the β -ketoaldehyde. Try to shoot for this intermediate. Show the mechanism to take intermediate A to the product.



Intermediate X